

Chemistry PPE Revision tasks

Task 2 - Trends on the periodic table

Resources

AQA Specification (available online):

5.1.2.1 The periodic table

5.1.2.4 Group 0, 5.1.2.5 Group 1, 5.1.2.6 Group 7

Collins Revision Guide

BBC Bitesize - <https://www.bbc.com/bitesize/guides/ztrxdxs/revision/1>

Periodic Table part 1 – Group 0, group 1 and group 7

1. What are the elements in Group 0 called?
2. What are the elements in Group 1 called?
3. What are the elements in Group 7 called?
4. What happens to the boiling point of elements in Group 0 as you go down the group?
5. Why are the elements in Group 0 so unreactive?
6. Why do all elements in Group 1 react in a similar way to each other?
7. What happens to the reactivity of the elements as you go down Group 1?
8. Write a word equation for the reaction between sodium and oxygen.
9. Why do all the elements in Group 7 react in a similar way to each other?
10. Halogens are diatomic. What does the word 'diatomic' mean?

11. What happens to the reactivity as you go down Group 7?
12. What happens to the melting point and boiling point as you go down Group 7?
13. Write a word equation for the reaction between lithium and chlorine.

Triple Science ONLY

Periodic Table part 2 – Transition metals

1. Where are transition metals found on the periodic table?
2. How do the melting points of transition metals compare to Group 1 metals?
3. How do the densities of transition metals compare to Group 1 metals?
4. How does the strength of transition metals compare to Group 1 metals?
5. Describe the differences between the reactions of the alkali metals and the reactions of transition metals.
6. State two typical properties of transition metals.
7. State one use of transition metals.
8. Explain why copper is used for plumbing.